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| Impurities In

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*Why does ice*  
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~~Waals - 4 types~~  
~~of Chemical~~  
~~Problem~~  
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~~Salts Hard~~  
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K how K depends  
on rxn

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Systems”. The  
Water Molecule:  
a Review. •

Water is a  
simple tri-  
atomic molecule,  
H<sub>2</sub>O. • Each O-  
H bond is highly  
polar, because

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of the high elec  
tronegativity of  
the oxygen (N,  
O, F, and Cl  
have high  
values). •bond  
angle of water =  
105°.

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Chapter 15 -  
Water and

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- 15.2

Homogeneous

Aqueous Systems

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Check - Page

501: 11 Answer

All ionic

compounds are

electrolytes

because they

dissociate into

ions and conduct

an electrical

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current in  
aqueous solution  
Water And  
or in the molten  
Aqueous  
state.

## Systems

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System. the  
chaotic movement  
of colloidal

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particles, caused by collision with particles of the solvent in which they are dispersed. a compound that conducts an electric current when it is in an aqueous solution or in the molten state; all ionic

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Compounds are  
electrolytes,  
but most  
covalent  
compounds are  
not.

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Chapter 15

"Water and

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"Aqueous Systems". use these activities to yourself

study the vocabulary and major concepts presented in

this chapter. heterogeneous mixture

containing particles that are small enough

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to remain  
dispersed in the  
solvent and do  
not separate on  
standing.

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Answers.

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speaking about  
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Answers, be sure  
to understand  
that Schooling  
is usually the

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crucial for an improved tomorrow, in addition to discovering won't only halt after a school bell rings. In which currently being reported, many of us provide various uncomplicated but helpful

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content and  
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created suited  
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to almost any  
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**Answers . . .**

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Water and

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Aqueous Systems.

surface tension.

surfactant.

aqueous

solution.

solvent. the

inward force or

pull that tends

to minimize the

surface ar... any

substance that

interferes with

hydrogen bonding

between wa... is

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SECTION 15.1

WATER AND ITS  
PROPERTIES

(pages 445–449)

This section  
describes the  
properties of  
water in the  
liquid and solid  
states and

# Where To Download

Chapter 15  
explains how  
hydrogen bonding  
affects the  
surface tension  
and vapor  
pressure of  
water. Water in  
the Liquid State  
(pages 445–447)  
1.

## **SECTION 15.1** **WATER AND ITS** **PROPERTIES**

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**(pages 445–449)**

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answers , source

:eur-

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may want to

stretch himself

once a worker

knows his

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efforts do not go unnoticed. For instance, if he knows his performance will be judged based on achievement of a target, he will work harder to achieve it.

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## Worksheet 15

### Answers

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Gravity. Created  
by. ... a

crystal of solid  
NaCl is placed  
into an aqueous

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NaCl solution.

No precipitate  
forms in the  
bottom of the

container. The

final solution

is: a.) super

saturated ...

Chapter 5: Soil

Water 57 Terms.

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OBJECTIVES:

-Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.

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this chapter.  
heterogeneous  
mixture  
containing  
particles that  
are small enough  
to remain

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dispersed in the solvent and do not separate on standing.

## **Chapter 15 Water Aqueous Systems Test B Answers**

CHAPTER 15 |

Aqueous

Equilibria:

Chemistry of the

Water World

15.1. Collect

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and Organize

Figure P15 shows  
.1 four lines to  
describe the

possible  
dependence of  
percent

ionization of  
acetic acid with  
concentration.

We are to choose  
the one that  
best represents  
the trend for

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this weak acid.  
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